

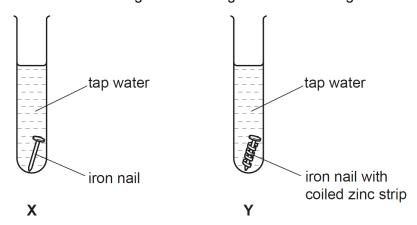
## **GCSE Chemistry B (Twenty First Century Science)**

J258/03 Breadth in chemistry (Higher Tier)

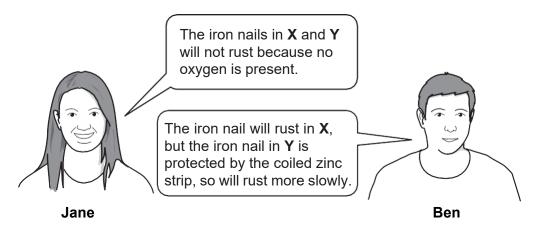
**Question Set 30** 

1 Iron is an important metal in the construction industry. The main disadvantage of iron is that it rusts.

Jane and Ben investigate the rusting of iron nails using two test tubes, **X** and **Y**.



(a) Jane and Ben discuss the investigation:



Do you agree with each person's comments?

Give one reason for each of your answers.

[2]

**(b)** When iron rusts, an iron(III) cation is made.

Fe → Fe<sup>3+</sup> + 3e<sup>-</sup>

(i) What is being oxidised in this reaction?

Put a (ring) around the correct formula.

Fe Fe<sup>3+</sup> 3e<sup>-</sup>

[1]

(ii) Give one reason for your answer to (b)(i).

[1]

(c)	Jane collects the rust. She dissolves the rust in hydrochloric acid and adds some sodium hydroxide solution.  She sees a brown precipitate.  What is the name of this brown precipitate?	
	Tick (√) <b>one</b> box.	
	Iron(III) chloride	
	Iron(II) hydroxide	
	Iron(III) hydroxide	
	Sodium chloride	

[1]

**Total Marks for Question Set 30: 5** 



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